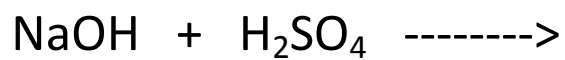


Titration Notes 2



Modified Titration Formula:

$$M_1V_1/C_1 = M_2V_2/C_2$$

M_1 = concentration of base

V_1 = volume of base

C_1 = coefficient of base

M_2 = concentration of acid (unknown)

V_2 = volume of acid

C_2 = coefficient of acid

Problems:

You titrated 25.0 mL of sulfuric acid with 52.33 mL of NaOH. The concentration of the base is 0.150M. What is the concentration of the acid?

You titrated 35.0 mL of phosphoric acid with 43.39 mL of $\text{Ca}(\text{OH})_2$. The concentration of the base is 0.200M. What is the concentration of the acid?

